

In the blank to the left, identify the type of reaction for each equation below. In the space below each equation, write a balanced chemical equation. For single replacement and double replacement reactions, assume that compounds are in aqueous solution.

_____ 1. Potassium and iodine react to form potassium iodide.

_____ 2. Carbonic acid and sodium hydroxide react to yield sodium carbonate and water.

_____ 3. Sodium and water react to form sodium hydroxide and hydrogen gas.

_____ 4. Magnesium bromate decomposes to form magnesium bromide and oxygen gas.

_____ 5. Carbon dioxide and water *are produced* when methanol burns in air

_____ 6. Zinc burns in air to form zinc oxide.

_____ 7. Copper (I) sulfate and lithium chloride react to form copper (I) chloride and lithium sulfate.

_____ 8. When heated, sodium carbonate forms sodium oxide and carbon dioxide.

_____ 9. Carbon dioxide and water react to form carbonic acid.

_____ 10. Aluminum and iron (II) acetate react to yield iron and aluminum acetate.

_____ 11. Cyclobutadiene (C₄H₄) burns in air

_____ 12. Calcium fluoride and silver (I) nitrate react.

_____ 13 Lithium burns in air.

_____ 14. Sulfuric acid and magnesium hydroxide react.

_____ 15. Benzene gas (C₆H₆) gas burns in air