

## CHAPTER 8

### THINGS TO KNOW

- ☐ From chapter 8
  - All the vocabulary and notes from the sections we've covered in chapter 7
  - Evidences of a chemical change
  - Diatomic molecules
  - Symbols used in chemical equations
  - General equations for the major types of chemical reactions
- ☐ Old from chapter 7
  - Memorize polyatomic ions (*the list is under the Worksheet tab on the class website*)
  - Memorize the formulas and names for the binary acids and oxyacids (*Worksheet tab on the class website*)
  - Memorize the charge/ oxidation number for groups 1, 2, and 13 - 17
  - Memorize the chemical formulas for ammonia, (NH<sub>3</sub>) and methane (CH<sub>4</sub>)

### THINGS TO KNOW HOW TO DO

- ☐ Write formulas for ionic compounds (*from Ch 7*)
- ☐ Write formulas for binary molecular compounds (*from Ch 7*)
- ☐ Read a chemical formula
  - 6Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> contain how many atoms or moles of Al? S? O?
- ☐ Balance chemical equations
- ☐ Read chemical equations (be able to use coefficients to show the ratio between elements)
- ☐ Complete chemical reactions given reactants
- ☐ Determine what type of reaction a chemical reaction is (from the 17 types of chemical reactions)
- ☐ Use the activity series for metals and nonmetals
  - To determine if single replacement reactions can occur
  - To determine if metals can replace hydrogen in water, steam, acids
  - To determine if oxides can form (if metals burn in air)
  - To determine if metal oxides can decompose

### THINGS YOU KNOW YOU NEED TO PRACTICE

- ☐
- ☐
- ☐
- ☐
- ☐