Dalton's Atomic Theory vs. Modern Atomic Theory

1. In the table below, use a different color highlighter or colored pencil/pen and draw a line from each postulate of Dalton's Atomic Theory to the one that most closely matches a postulate from the Modern Atomic Theory. You will need 5 colors. You will use one from the Modern Atomic Theory twice. CAUTION: The same number of each theory may not match!!
2. In the Compare & Contrast column, color the oval with the color you used to match the first postulate from Dalton's Atomic Theory and then explain how the theories are the same, different or BOTH! Do this with all 5 postulates of Dalton's Atomic Theory.

Dalton's Atomic Theory	Modern Atomic Theory	Compare & Contrast
1. All matter is composed of extremely small particles called atoms.	1. All matter is composed of atoms.	
2. Atoms of an element are identical in size, mass, and other properties; atoms of different elements differ in size, mass, and other properties.	2. Atoms of the same element have the same chemical properties; atoms of different element have different chemical properties.	
3. Atoms cannot be subdivided, created, or destroyed.	3. Atoms of an element have a characteristic average mass which is unique to that element.	
4. Atoms of different elements combine in simple whole-number ratios to form chemical compounds.	4. Atoms cannot be subdivided, created, or destroyed in ordinary chemical reactions. However, these changes can occur in nuclear reactions.	
5. In chemical reactions, atoms are combined, separated, or rearranged.	5. Compounds are formed when atoms of two or more elements unite, each atom losing its own characteristic properties