CHEMISTRY
BALANCING EQUATIONS WORKSHEET IV NAME_____

ANSWER THREE PARTS: 1) Identify the type of each equation in the blank to the left. **2)** Write the rule that applies to each in the blank beside the word equation. **3)** Complete and balance the chemical equations in the space below each.

_____ 1. Burning of magnesium in air _____

_____ 2. Decomposition of copper (II) carbonate _____

_____ 3. Neutralization of magnesium hydroxide and phosphoric acid _____

_____ 4. Addition of silver (I) to nitric acid _____

_____ 5. Decomposition of aluminum chlorate _____

_____ 6. Addition of calcium oxide to water _____

_____ 7. Addition of chlorine gas to ammonium iodide _____

_____ 8. Combustion of butane gas (C_4H_{10}) _____

_____ 9. Reaction between mercury (II) sulfate and ammonium nitrate _____

_____ 10. Addition of aluminum to hydrochloric acid _____

_____ 11. Reaction between dinitrogen pentoxide and water _____

_____ 12. Decomposition of iron (III) hydroxide _____

_____ 13. Addition of aluminum to copper (II) sulfate _____

_____ 14. Burning propane (C_3H_8) gas in air _____

_____ 15. Reaction between sodium sulfate and barium nitrate _____

_____ 16. Reaction between aluminum and bromine gas _____

_____ 17. Decomposition of lithium oxide _____

_____ 18. Addition of zinc metal to steam _____

_____ 19. Lead (II) nitrate reacts with potassium iodide _____

_____ 20. Glucose ($C_6H_{12}O_6$) burns in air _____