

IDENTIFY EACH OF THE TYPES OF EQUATIONS BELOW IN THE BLANK TO THE LEFT, THEN COMPLETE AND BALANCE THE FOLLOWING CHEMICAL EQUATIONS:

- _____ 1. Potassium burns in air to produce potassium oxide
- _____ 2. Addition of calcium oxide to water to produce calcium hydroxide
- _____ 3. When heated copper (II) carbonate decomposes to produce copper (II) oxide and carbon dioxide
- _____ 4. Reaction between barium hydroxide and hydrofluoric acid to produce barium fluoride and water
- _____ 5. Cesium reacts with water to produce cesium hydroxide and hydrogen gas
- _____ 6. Chlorine gas reacts with magnesium bromide to produce magnesium chloride and bromine gas
- _____ 7. Acetylene gas(C_2H_2) gas burns in air to produce carbon dioxide and water
- _____ 8. Lead (II) phosphate reacts with ammonium sulfide to produce lead (II) sulfide and ammonium phosphate
- _____ 9. Carbonic acid is produced when carbon dioxide reacts with water
- _____ 10. Aluminum hydroxide decomposes to produce aluminum oxide and water
- _____ 11. Carbon dioxide and water are produced when pentane (C_5H_{12}) gas burns in air
- _____ 12. Lithium chloride and iron (II) nitrate react to produce lithium nitrate and iron (II) chloride
- _____ 13. Nickel (I) oxide decomposes to produce nickel and oxygen gas
- _____ 14. Potassium and aluminum permanganate react to produce aluminum and potassium permanganate
- _____ 15. Tetraphosphorus decoxide reacts with water to produce phosphoric acid