CHEMISTRY
BALANCING EQUATIONS WORKSHEET Ia NAME_____

USE THE DEFINITIONS OF <u>SYNTHESIS</u>, <u>DECOMPOSITION</u>, <u>SINGLE DISPLACEMENT</u>, <u>DOUBLE DISPLACEMENT</u>, AND <u>COMBUSTION</u> REACTIONS TO IDENTIFY EACH TYPE OF REACTION BELOW AND WRITE THE TYPE IN THE BLANK TO THE LEFT. NEXT, UNDER THE WORD EQUATION, WRITE A BALANCED CHEMICAL EQUATION FOR EACH REACTION.

_____ 1. Lithium and fluorine react to form lithium fluoride.

_____ 2. Nitric acid and magnesium hydroxide react to yield magnesium nitrate and water.

_____ 3. Iron (III) oxide and carbon monoxide react to form iron and carbon dioxide gas.

_____ 4. Ammonium chloride and barium hydroxide react to yield barium chloride and ammonia and water.

_____ 5. Copper and oxygen react to yield copper (I) oxide.

_____ 6. Sodium sulfate and barium chloride react to form sodium chloride and barium sulfate.

_____ 7. Sodium and water react to yield sodium hydroxide and hydrogen.

_____ 8. Hydrogen and chlorine react to form hydrochloric acid.

_____ 9. Magnesium and water react to yield magnesium hydroxide and hydrogen.

_____ 10. Iron and chlorine react to form iron (III) chloride.

_____ 11. Sodium chloride and silver (I) nitrate react to yield silver (I) chloride and sodium nitrate.

_____ 12. Magnesium sulfate and ammonium chloride react to yield magnesium chloride and ammonium sulfate.

_____ 13. Sulfuric acid and potassium hydroxide react to yield potassium sulfate and water.

_____ 14. Sulfur dioxide and oxygen react to form sulfur trioxide.

_____ 15. Aluminum and iron (II) nitrate react to yield iron and aluminum nitrate.